

ELECTROCHEMICAL CELL REACTION PRACTICE

Given the following reactants, write balanced equations for each possible combination of neutral metal with a compound. Determine if there would be a reaction. If there is a reaction, write the half reactions, and indicate which metal is oxidized and which is reduced. Then write the cell notation for each electrochemical cell, and calculate the cell voltage.

Given: solid Al, Ag, Mg, Zn, and Cu, and 0.1M solutions of: Al^{3+} , Ag^+ , Mg^{2+} , Zn^{2+} , and Cu^{2+}

→ oxidized = _____ reduced = _____

→ cell notation: _____

→ $E^\circ_{\text{cell}} =$ _____

→ oxidized = _____ reduced = _____

→ cell notation: _____

→ $E^\circ_{\text{cell}} =$ _____

→ oxidized = _____ reduced = _____

→ cell notation: _____

→ $E^\circ_{\text{cell}} =$ _____

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